

# READING ASSIGNMENT FOR NOVEMBER 15

## SECTIONS 11.7 AND 11.8

### 11.7 - Entropy and the Second Law of Thermodynamics

- Irreversible process - happens in one direction only.
- The book's definition of entropy is fairly unusual, but a good one. Entropy is usually defined as the amount of "disorder". I think the book's definition makes it more precise.
- Second Law - only applies to isolated systems.
- An increase in thermal energy will also come along with an increase in entropy.
- I might bust out a chemistry equation to explain where the maximum efficiency and *COP* equations come from.

### 11.8 - Systems, Energy, and Entropy

- This section mostly wraps things up for you. Enjoy.